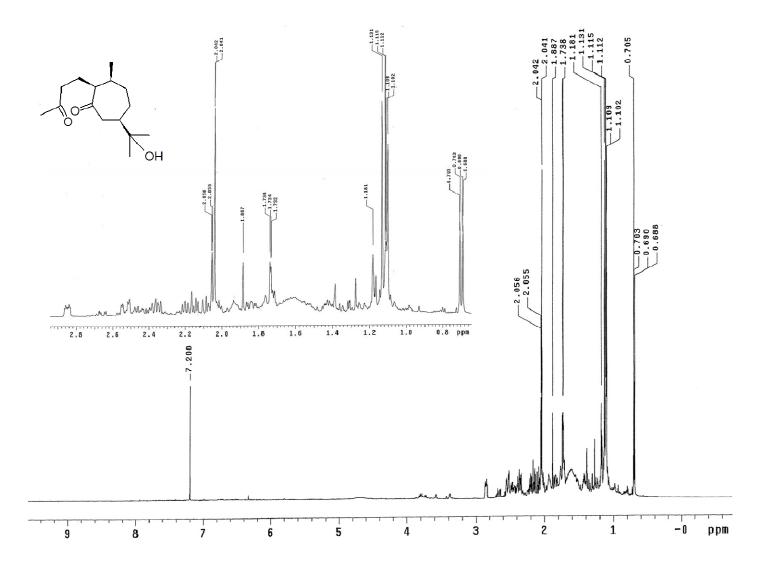
Supplementary Materials

Figure S1. ¹H-NMR (500 MHz, CDCl₃) spectrum of the new compound **1**.



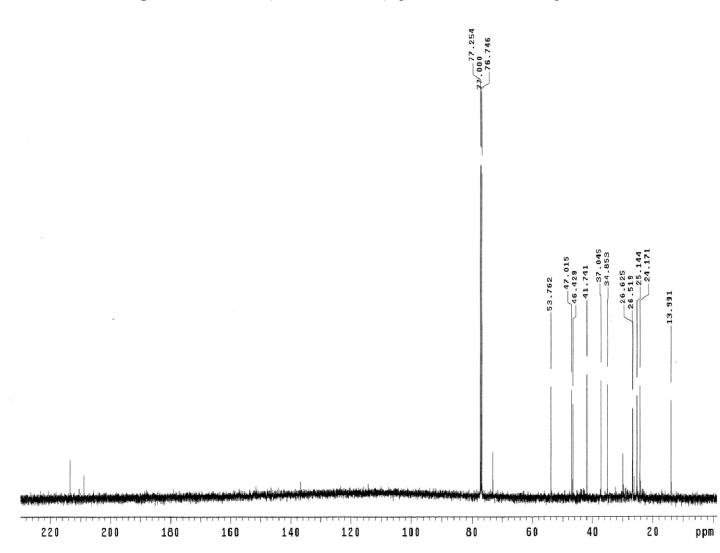


Figure S2. ¹³C-NMR (126 MHz, CDCl₃) spectrum of the new compound 1.

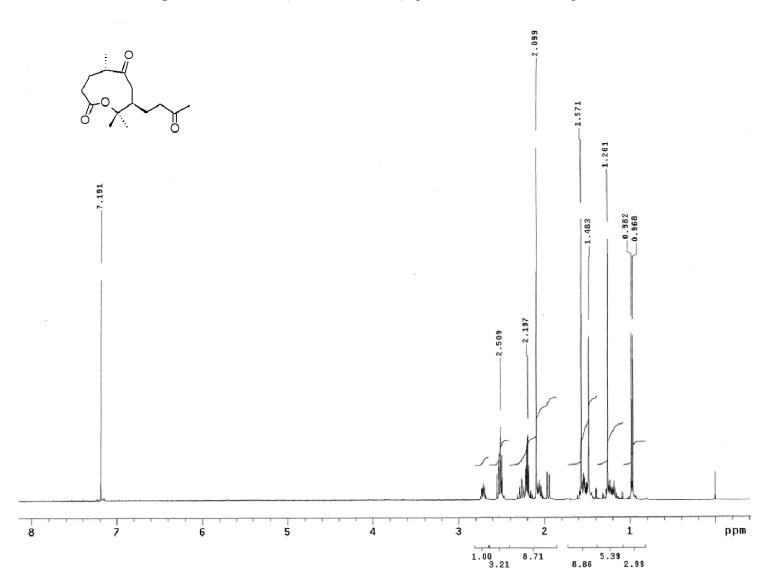


Figure S3. ¹H-NMR (500 MHz, CDCl₃) spectrum of the new compound **2**.

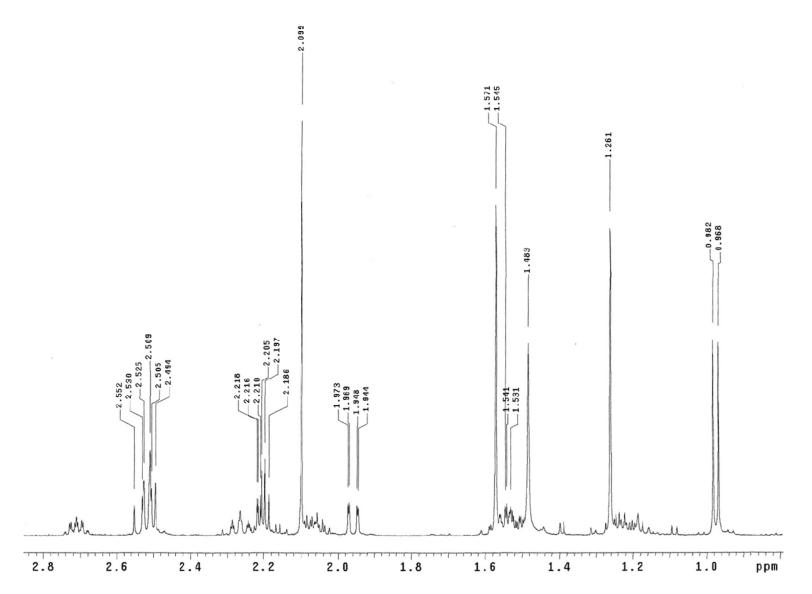
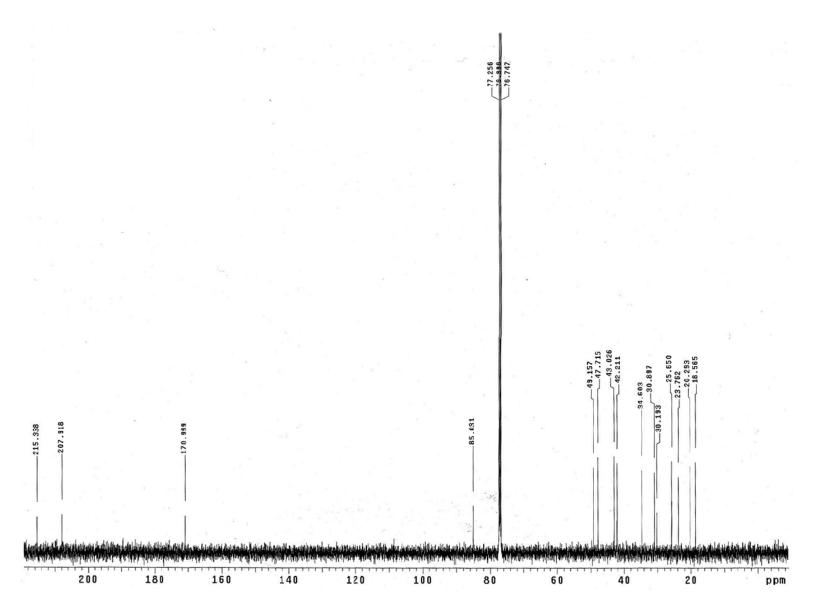


Figure S4. Amplified ¹H-NMR (500 MHz, CDCl₃) spectrum of the new compound **2**.

Figure S5. ¹³C-NMR (126 MHz, CDCl₃) spectrum of the new compound **2**.



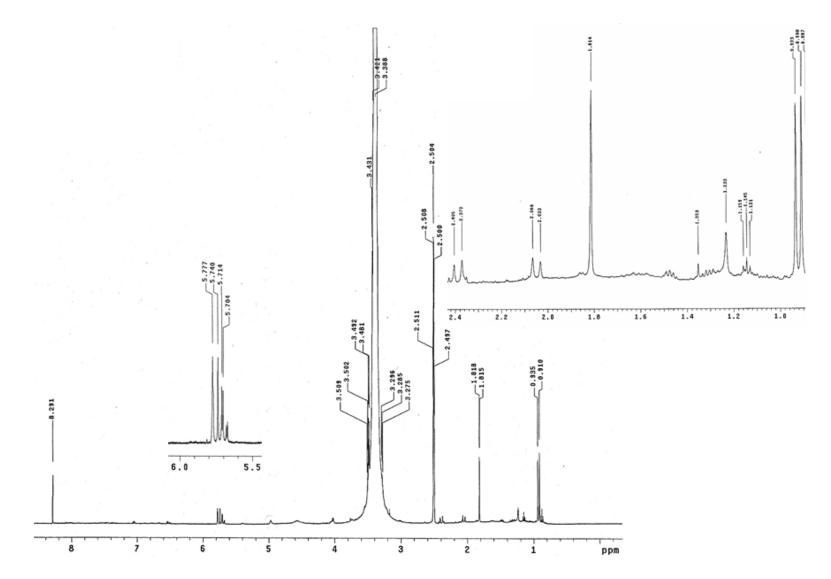


Figure S6. ¹H-NMR (500 MHz, DMSO- d_6) spectrum of the new compound **3**.

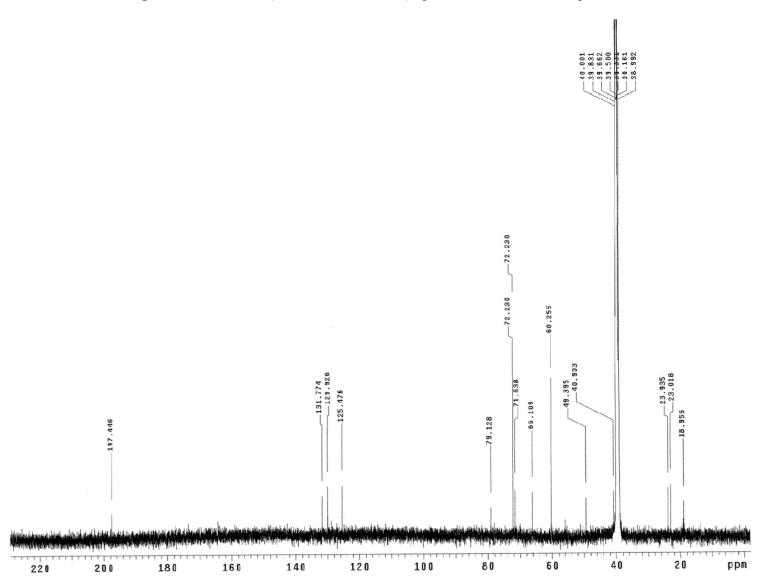
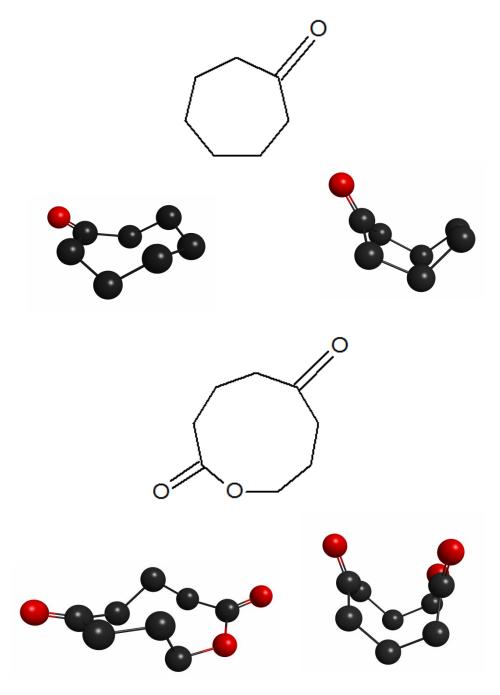


Figure S7. ¹³C-NMR (126 MHz, DMSO- d_6) spectrum of the new compound **3**.

Computational Details

Calculations have been carried out using PC GAMESS package at the B3LYP/STO-3G** level and wxMacMolPlt software for structure visualization [1–6]. Optimization algorithm was based on the Quadratic Approximation (QA) and the threshold gradient value was 10^{-5} a.u. [7]. Vibrational analysis showed all real frequencies asserting the stationary points as minima [7]. The NBO donor-acceptor pairs were checked and steric energies were calculated [8–12].

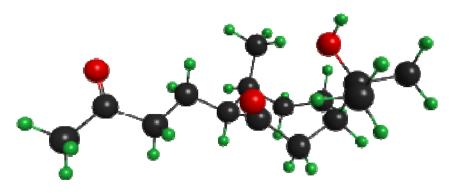
Figure S8. Conformations found in the Cambridge Structural Database.



7-membered ring			
Conformer	ΔG ^o (298.15 K)	ΔE_{ZPE}	ΔΕ
71	0.0	0.0	0.0
72	0.0	0.2	0.7
73	0.4	1.2	2.0
74	-0.6	1.1	2.4
9-membered ring			
Conformer	ΔG ^o (298.15 K)	ΔE_{ZPE}	ΔΕ
91	0.0	0.0	0.0
92	4.3	4.0	3.7
93	14.4	15.3	15.9
94	22.6	22.3	22.0

Table S1. Relative Energies (kcal.mol⁻¹)—Conformation.

Figure S9. Optimized structures for the 7-membered ring compound.



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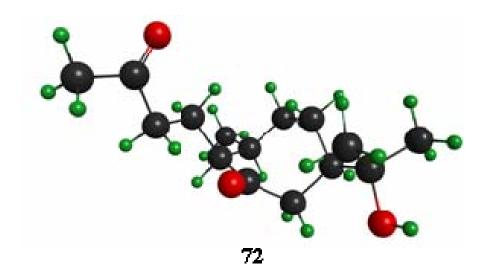


Figure S9. Cont.

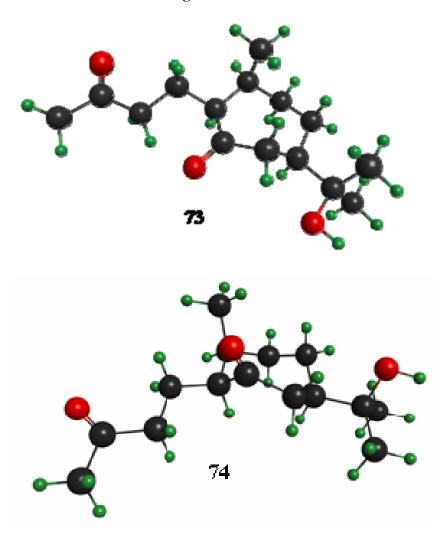
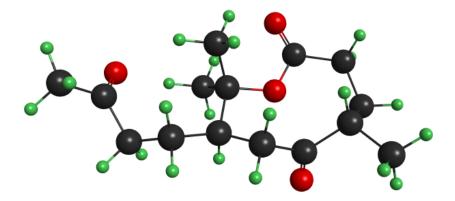
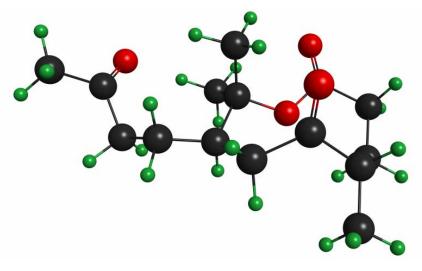
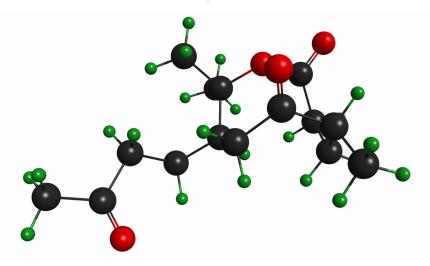


Figure S10. Optimized structures for the 9-membered ring compound.

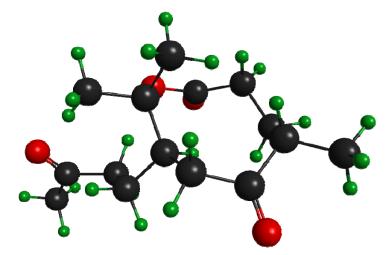








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References

- 1. PC GAMESS. Available online: http://classic.chem.msu.su/gran/gamess/index.html (accessed on 31 November 2011).
- Becke, A.D. Density-functional thermochemistry. III. The role of exact exchange. J. Chem. Phys. 1993, 98, 5648–5652.
- 3. Hehre, W.J.; Stewart, R.F.; Pople, J.A. Self-consistent molecular-orbital methods. I. Use of gaussian expansions of slater-type atomic orbitals. *J. Chem. Phys.* **1969**, *51*, 2657–2664.
- Hehre, W.J.; Ditchfield, R.; Stewart, R.F.; Pople, J.A. Self-consistent molecular orbital methods. IV. Use of gaussian expansions of slater-type orbitals. Extension to second-row molecules. *J. Chem. Phys.* 1970, 52, 2769–2773.
- Collins, J.B.; Schleyer, P. von. R.; Binkley, J.S.; Pople, J.A. Self-consistent molecular orbital methods. XVII. Geometries and binding energies of second-row molecules. A comparison of three basis sets. J. Chem. Phys. 1976, 64, 5142–5151.
- Bode, B.M.; Gordon, M.S. Macmolplt: A graphical user interface for GAMESS. J. Mol. Graphics Mod. 1998, 16, 133–138.
- 7. Jensen, F. Locating transition structures by mode following: A comparison of six methods on the Ar₈ Lennard-Jones potential. *J. Chem. Phys.* **1995**, *102*, 6706–6718.
- 8. Foster, J.P.; Weinhold, F. Natural hybrid orbitals. J. Am. Chem. Soc. 1980, 102, 7211–7218.
- 9. Reed, A.E.; Weinstock, R.B.; Weinhold, F. Natural population analysis. J. Chem. Phys. 1985, 83, 735–746.
- 10. Reed, A.E.; Weinhold, F. Natural localized molecular orbitals. J. Chem. Phys. 1985, 83, 1736-1740.
- 11. Badenhoop, J.K.; Weinhold, F. Natural bond analysis of steric interactions. J. Chem. Phys. 1997, 107, 5406–5421.
- 12. Badenhoop, J.K.; Weinhold, F. Natural steric analysis of internal rotation barriers. *Int. J. Quantum Chem.* **1999**, *72*, 269–280.